AP Chemistry – Equation Balancing

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| **Focus and Review** | Students will review solubility rules and predict solubility of several compounds.  State whether the following compounds are soluble or insoluble:   1. NaBr 2. AgCl 3. MgSO4 4. FeSO4 5. Ca(NO3)2 6. CaCO3 |
| **Objective/Goal** | 4.01 Analyze the various types of common chemical reactions  Precipitation Reactions |
| **Instructional Presentation** | Students are presented with what solubility rules are and predict reactions.  Before we even go to the lab, students will be presented with the list of reactions they will investigate. They make predictions ahead of time. In the interaction, they will be checking their predictions.  We go to the computer lab (my AP students do not have laptops) to access the website  <http://www.sascurriculumpathways.com/ProductEntrance/Navigation/navigator.jsp?node=54>  Click on the InterActivity. |
| **Guided Practice** | We go over the first two examples together to make sure students know how to manipulate the software, balance it, etc. They view a video that shows a real-world example of the reaction they’ve set up. This allows them to see reactions of more dangerous chemicals (strong oxidizers like AgNO3). |
| **Independent Practice** | Students complete the activity on their own, carrying out each experiment, writing down the results and reporting what they find. |
| **Closure/Reflection** | Students will see patterns from balancing reactions and predicting soluble products.  Students can go back and generalize solubility rules simply from what we’ve seen demonstrated. |